

## EFFECT OF Mg (II) METAL ION ON ANTIOXIDANT ACTIVITIES OF ETHANOL EXTRACTS OF RAMBUTAN PEEL (*Nephelium lappaceum*)

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**Abstrak.** Antosianin adalah salah satu pewarna alami yang banyak digunakan dalam bahan pangan, karena warnanya yang dapat beragam mulai dari merah, biru dan ungu. Penyimpanan antosianin dapat mengakibatkan antosianin terdegradasi, baik adanya interaksi dengan suhu, cahaya, dan oksigen sehingga dapat mempengaruhi stabilitas warna antosianin sendiri, bahkan dapat mengurangi kemampuan antosianin sebagai antioksidan. Permasalahan tersebut dapat diatasi dengan kopigmentasi, yaitu dengan mereaksikan antosianin dengan ion logam. Pada penelitian ini logam yang digunakan adalah  $Mg^{2+}$ . Hasil penelitian memperlihatkan bahwa ekstrak etanol kulit buah rambutan diperoleh ekstrak kental sebanyak 23,824 gram dengan kadar total antosianin 709,3866 mg/L dengan aktivitas antioksidan 0,0046  $\mu\text{g/mL}$ , ekstrak antosianin dengan penambahan  $Mg^{2+}$  50, 100 dan 150 ppm memiliki nilai  $IC_{50}$  berturut-turut sebesar 0,0061; 0,0048 dan 0,0067  $\mu\text{g/mL}$ . Hal ini menunjukkan bahwa penambahan ion logam  $Mg^{2+}$  memperlihatkan perbedaan aktivitas antioksidan antosianin yang tidak signifikan.

Kata kunci: antioksidan, antosianin, magnesium, rambutan

**Abstract.** Anthocyanin is the one of natural dyes that often used in foodstuff, because anthocyanin can make a more color like red, blue and purple. Anthocyanin storage can make anthocyanin is damage, because the interaction with temperature, light, and oxygen show that can given effect about the stability anthocyanin color, and decrease the anthocyanin as an antioxidant. That problem can be solved with copigmentation, that is an anthocyanin reaction with metal ion. In this study can showed that ethanol extract from Rambutan peel and we obtained 23.824 gram with total anthocyanin content of 709.3866 mg/L with antioxidant activity of 0.0046  $\mu\text{g/mL}$ , and then the extract of anthocyanin with metal ion  $Mg^{2+}$  in concentration 50; 100 and 150 ppm had  $IC_{50}$  are 0.0061; 0.0048 and 0.0067  $\mu\text{g/mL}$ . This conclusion from the added of metal ion  $Mg^{2+}$  showed that has not show much different antioxidant activity.

Keywords: antioxidant, anthocyanin, magnesium, rambutan.

## INTRODUCTION

Indonesia is an agribusiness country with the largest income centered on agriculture and plantations with abundant produce. One of them is rambutan fruit, with an average annual production from 2009-2014 of 733,105.5 tons (Ministry of Agriculture, 2015). This plant originated from Indonesia, and began to spread to other tropical countries such as the Philippines and countries in Latin America (Mahisworo et al., 1991).

Rambutan is known to have various pharmacological activities such as antidiabetic, antihypercholesterol, antimicrobial, antioxidant, antihyperuricemia, and anticancer (Sadino, 2017). In 2014 Prasetyo used rambutan rind as a textile natural dye, because it contained anthocyanin pigments.

According to Francis (1982), anthocyanin is a polar compound, so it can be extracted with polar solvents, such as water, ethanol, and methanol. Solvents that are often used to extract anthocyanins are methanol or ethanol which is acidified with HCl. However, due to the toxic nature of methanol, ethanol is usually used in food systems acidified with HCl. Based on the research conducted by Syarifuddin (2011), Moulana et al. (2012), Hutapea et al. (2014) and Wulandari (2014) good solvents used to extract anthocyanin are ethanol solvents, because the total anthocyanin content produced is greater than that of methanol.

Anthocyanin pigments have biological activities that are beneficial to the health of the body. But along with the length of storage, anthocyanins can experience damage due to environmental influences during storage. One way to

overcome this problem is through a process with anthocyanin copigmentation with Mg (II) metal (Brouillard, 1982).

Kondo et al. (1992) prove that the complex between Mg<sup>2+</sup> and anthocyanin produces a more stable pigment complex. Magnesium (Mg) plays a role in the smooth operation of various enzymes. Magnesium is needed by the body to produce 300 types of enzymes, sending messages through the nervous system, making the muscles stay flexible and relaxed and maintaining bone and tooth strength (Budiasih, 2009).

Based on the description above, further research is needed on the effect of adding metal ions Mg (II) to the anthocyanin activity of the ethanol extract of rambutan peel.

In this study anthocyanin extraction from rambutan peel using ethanol solvent was extracted, then extracted with Mg metal and tested for antioxidant activity before and after copigmentation using the DPPH method.

Indonesia merupakan negara agribisnis dengan penghasilan terbesarnya terpusat pada pertanian dan perkebunan dengan hasil buminya yang melimpah. Salah satunya adalah buah rambutan, dengan rata-rata produksi per tahunnya dari 2009-2014 sebanyak 733.105,5 ton (Kementerian Pertanian, 2015). Tanaman ini asli berasal dari Indonesia, dan mulai menyebar ke negara tropis lainnya seperti Filipina dan negara-negara di Amerika Latin (Mahisworo dkk., 1991).

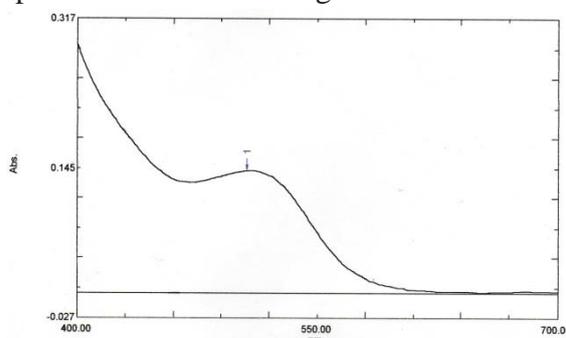
## MATERIAL AND METHOD

This research was carried out by extracting anthocyanin from rambutan

peel, then characterized by UV-Vis spectrophotometer (spectrum and total anthocyanin content) and Infrared Spectroscopy (functional group analysis), and determining anthocyanin activity by DPPH method.

## RESULT AND DISCUSSION

The crude extract obtained from extracting 120 grams of rambutan fruit skin was 23.824 grams. The extract obtained was analyzed by a UV-Vis spectrophotometer obtained by the spectrum as shown in Figure 1.



**Figure 1.** The maximum wavelength ( $\lambda_{maks}$ ) of ethanol extract from rambutan fruit skin

The maximum wavelength of ethanol extract of rambutan peel obtained is 506 nm. Harborne (1987), states that the maximum absorbance achieved by anthocyanins is at wavelengths of 490-550 nm. So the results obtained are in

accordance with the characteristics of the maximum anthocyanin wavelength and it can be said that rambutan peel extract contains anthocyanin.

Determination of anthocyanin levels was calculated using the pH difference method at pH 1 and pH 4.5. Based on the results of the calculation, the total anthocyanin level was 709.3866 mg / L.

$$A = (A_{\lambda_{maks}} - A_{700})_{pH 1} - (A_{\lambda_{maks}} - A_{700})_{pH 4,5}$$

$$A = (0,142 - 0,002) - (0,06 - 0,005)$$

$$A = 0,085$$

Total content of anthocyanin (mg/L)

$$\begin{aligned} & \frac{A \times BM \times FP \times 1000}{\epsilon \times l} \\ & = \frac{0,085 \times 449,2 \text{ g/mol} \times 500 \times 1000}{26900 \text{ L/mol.cm} \times 1 \text{ cm}} \\ & = 709,3866 \text{ mg/L} \end{aligned}$$

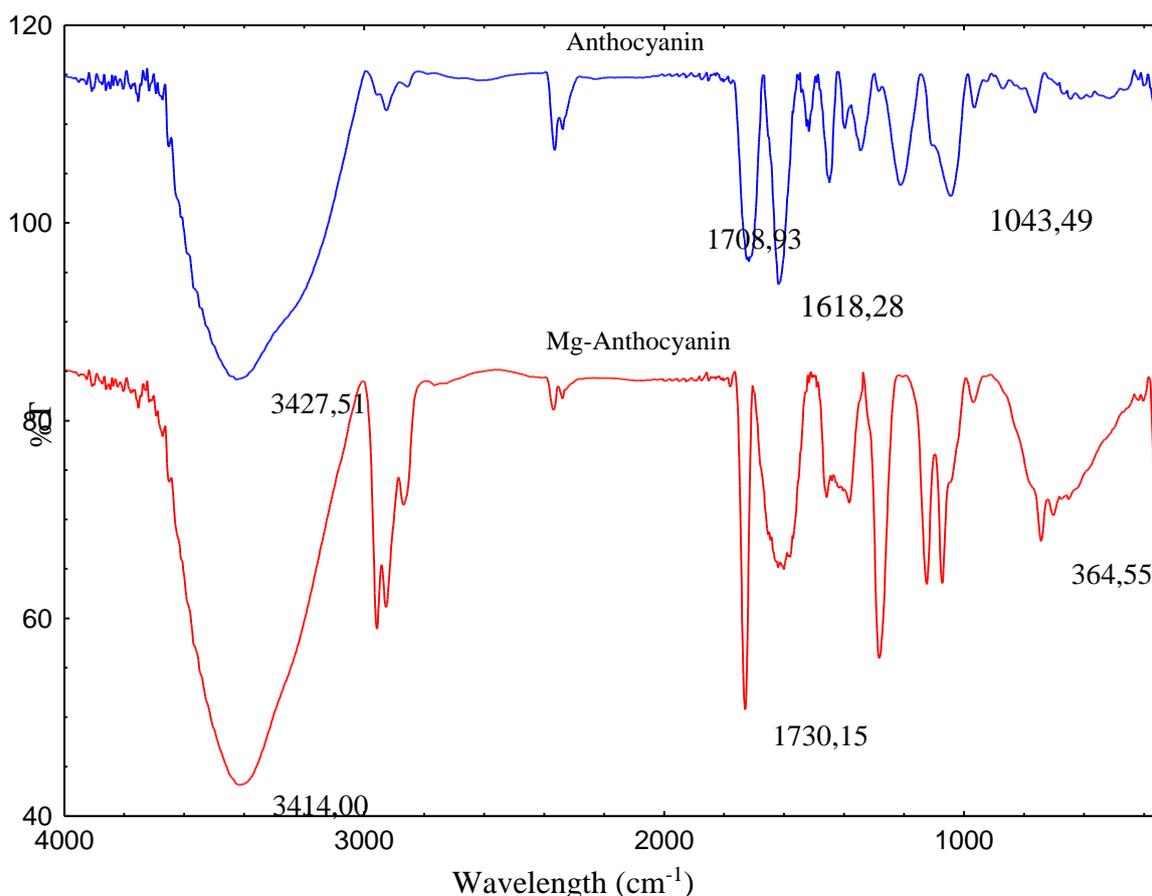
Where BM = molecule mass of cyanidin-3-glucoside (449,2 g/mol)

DF = dilution factor

$\epsilon$  = absorptivity molar cyanidin-3- glucoside (26.900 L/mol.cm)

L = cuvette thick (1 cm)

Analysis with infrared spectroscopy obtained the spectrum and wavelength before and after the addition of metal ion  $Mg^{2+}$  can be seen in Figure 2.

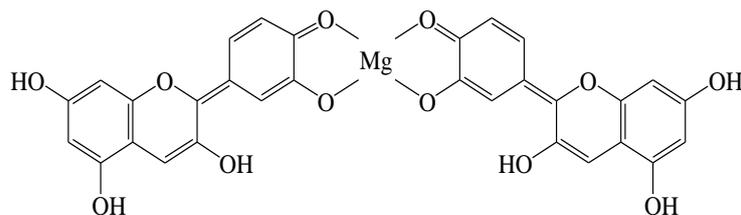


**Figure 2.** IR spectrum of anthocyanin extract before and after addition of metal (Kopigmentation)

The results of the FTIR spectrum (Figure 2) show a sharp uptake in the wave number area 3427.51 cm<sup>-1</sup> which indicates the presence of the -OH functional group supported by the appearance of absorption in the wave number region 1043.49 cm<sup>-1</sup> for the C-O alcohol functional group. Wave numbers also appear with values of 1618.28 cm<sup>-1</sup> and 1517.98 cm<sup>-1</sup> with sharp absorption, which indicates the presence of a functional group C = C aromatic which is reinforced by absorption at a wave number 763.81 cm<sup>-1</sup> for bond absorption CH aromatic and the absorption of C = O bonds is shown at wave number 1708.93 cm<sup>-1</sup>.

Based on FTIR analysis of the rambutan anthocyanin extract before and after the addition of metals showed that there were changes in wave numbers especially in OH groups with wave number values before the addition of metal ions of 3427.51 cm<sup>-1</sup> and after the addition of ions metal Mg<sup>2+</sup> changed to wave number 3414.00 cm<sup>-1</sup>.

The shift in wave number indicates that a metal ion binding process occurs with rambutan anthocyanin extract. This is in accordance with the statement of Miguel (2011), that anthocyanins and metals can be bound to hydroxyl groups.



**Figure 3.** Estimated structure of metal complexes with anthocyanins (Ahmed et al., 2013).

Besides the wave number at 364.55 cm<sup>-1</sup>, the addition of metal ion Mg<sup>2+</sup>, shows the M-O bond, indicating the formation of the Mg-Anthocyanin complex (Ahmed and BenGuzzi, 2008).

The results of the analysis of the antioxidant activity test using the DPPH

method, the absorbance of extracts and DPPH were measured by a spectrophotometer at the 515 nm wave length. Measurement of anthocyanin antioxidant activity with Vitamin C as a comparison can be seen in Table 1.

**Table 1.** IC<sub>50</sub> values and the level of antioxidant strength of ethanolic extract of rambutan fruit skin and Mg-anthocyanin extract and comparison (Positive Control)

Sample/ comparison	IC <sub>50</sub>	Intensity
Sample		
Anthocyanin	0,0046 µg/mL	Very Strong
Mg-Anthocyanin 50 ppm	0,0061 µg/mL	Very Strong
Mg-Anthocyanin 100 ppm	0,0048 µg/mL	Very Strong
Mg- Anthocyanin 150 ppm	0,0067 µg/mL	Very Strong
Comparison (Positive Control)		
Ascorbic Acid	0,2349 µg/mL	Very Strong

From these data it can be concluded that the anthocyanin of rambutan peels has a stronger antioxidant activity than the positive control of ascorbic acid. IC<sub>50</sub> value of rambutan skin extract was obtained 0.0046 µg/mL, the most powerful IC<sub>50</sub> Mg-anthocyanin value was 100 ppm. These results are obtained based on linear regression calculations from the absorbance curve of the sample. Based

on Table 1, it can be seen that anthocyanin without the addition of metals compared to the addition of metal ions Mg<sup>2+</sup> has antioxidant activity that is not significantly different. This is due to the addition of metals making anthocyanins more stable, stable compounds are increasingly difficult to react with other compounds (Takeda et al., (1994); Kondo et al. (1992).

## CONCLUSION

Based on the results of the research conducted, it was shown that the addition of metal ions Mg (II) showed insignificant differences in anthocyanin antioxidant activity.

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